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Section 10.3 – Percent Composition and Chemical Formulas. The percent by mass (percent composition) of an element in a compound is the number of

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grams of the element divided by the mass in grams of the compound multiplied by 100%. % mass of element = $\frac{\text{mass of element}}{\text{mass of compound}} \times 100$.

Chapter 10 - Chemical Quantities

CHAPTER 10: Chemical Quantities

BASICS: • The basic unit that is used to

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determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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the SI unit representing 6.02×10^{23}
representative particles of a substance.
the temperature and pressure at which
one mole of gas occupies a volume of

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22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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10.1 kg 103 g 368 Chapter 10 Chemical Calculations and Chemical Equations.
10.1 Equation Stoichiometry 369 The ratio of moles of P_4O_{10} to moles of P (which came from the subscripts in the chemical formula, P_4O_{10}) provided the key conversion factor that allowed us to convert

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Chapter 10 Chemical Calculations and equations

In this lesson, we'll explore what an enzyme is, what an enzymatic reaction is and what happens to the rate of a reaction if you add more enzymes

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key.

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Chemical Quantities THE MOLE AND
QUANTIFYING MATTER 10.1 The Mole: A
Measurement of Matter Essential
Understanding The mole represents a
large number of very small particles.

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Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

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Chapter 10 Chemical Quantities 91.

SECTION 10.1 THE MOLE: A

MEASUREMENT OF MATTER (pages

287-296) This section defines the mole

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and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287-289)

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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Chapter 10 Chemical Quantities 2.06%
32.69% 65.25% Acid X 2.46% 39.06%
58.48% Acid Y Use the circle graphs
above to answer the following questions.
1. 2. 3. What is the percent of carbon in
50.0 g of cyclohexane? Calculate the
mass of hydrogen in 150.0 g of ethene.
Using the circle graphs, explain why
percent composi

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worksheet Objectives: To enable

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